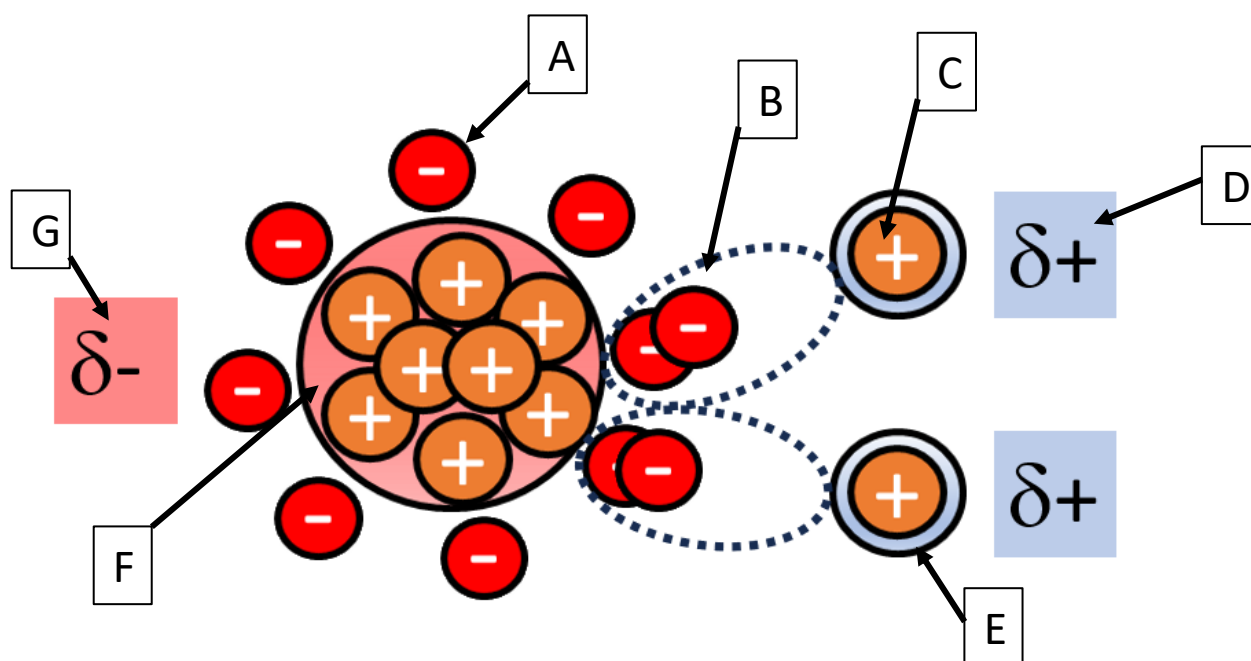


adhesion	the molecular force of attraction in the area of contact between unlike bodies that acts to hold them together
anion	a negatively charged ion
atom	the smallest component of an element having the chemical properties of the element
atomic mass	The mass of a given atom or molecule, expressed in atomic mass units (amu).
attract	to exert a force on (a body) that tends to cause an approach or oppose a separation
coefficient	a number or quantity placed before and multiplying another quantity, as 2 in the expression 2 H ₂ O.
cation	a positively charged ion
cohesion	The force of attraction that holds molecules of a given substance together
condense	when water vapour collects together to form liquid water as it cools
covalent bond	the bond formed by the sharing of a pair of electrons by two atoms
crystal	a piece of solid substance, such as table salt, with a regular shape due to the regular internal structure of its atoms, ions, or molecules
dissolve	of something solid : to mix with a liquid and become part of the liquid
electron	An elementary particle with a negative charge and a very small mass . Electrons are normally found in orbits around the nucleus of an atom. The chemical reactions that an atom undergoes depend primarily on the electrons in the outermost orbits (the valence electrons) .
filter paper	a porous paper used for filtering liquids
filtrate	the liquid that runs through a filter paper
hydrogen bond	the bond formed when the negative side of the water molecule attracts the positive side of another molecule
hydrophilic	having an affinity for water; loves water ☺
hydrophobic	having little or no affinity for water; hates water ☹
ion	particles (atoms or molecules) that have a net positive or negative charge
ionic bond	the electrostatic bond between two ions formed through the transfer of one or more electrons
molecule	two or more atoms held together by chemical bonds
negative ion	an electrically charged atom or group of atoms formed by the gain of one or more electrons; anion
net charge	In an ordinary atom, the number of protons equals the number of electrons, so the atom normally has no net electric charge. An atom becomes negatively charged if it gains extra electrons, and it becomes positively charged if it loses electrons; atoms with a _____ are called ions.
neutron	An elementary particle without an electrical charge ; one of the building blocks of the nucleus of the atom. A neutron has about the same mass as a proton .
opaque	light does not pass through this type of material
polar	a molecule in which there is an uneven distribution of electrons; one side of the molecule is slightly negative (δ-) the other side slightly positive (δ+); δ+ = "delta positive"
positive ion	an electrically charged atom or group of atoms formed by the loss of one or more electrons. cation
precipitate	to separate as a solid from a solution
product	a substance obtained from another substance through chemical change.
proton	An elementary particle with a positive charge , found in atomic nuclei in numbers equal to the atomic number of the element

reactant	a substance that participates in a chemical reaction, esp a substance that is present at the start of the reaction
repel	to push away
residue	the solid material that stays in the filter after the solution goes through
saturated	a solution that contains the maximum amount of solute for a given quantity of solvent at constant temperature and pressure
solubility	the amount of solute that dissolves in a given quantity of a solvent at a specified temperature and pressure to produce a saturated solution
solute	the dissolved particles in a solution
solution	a mixture in which particles of one or more substances (the solute) are distributed uniformly throughout another substance (the solvent)
solvent	a substance that dissolves another to form a solution
subscript	written below - as in H ₂ O or O ₂
surface tension	a property of liquids caused by intermolecular forces near the surface leading to the apparent presence of a surface film
suspension	a mixture in which small particles of a substance are dispersed throughout a gas or liquid.
transparent	light can pass through this type of material



Label Figure 1

Figure 1

A.....

E.....

B.....

F.....

C.....

G.....

D.....